

<b>Term</b>	<b>Definition</b>
Amphoteric	A substance that can <b>ACT LIKE AN ACID OR A BASE</b>
Arrhenius acid	a substance whose water (aqueous) solution contains or yields <b>HYDRONIUM IONS (H<sup>+</sup> IONS)</b> as the <b>ONLY POSITIVE ION</b> in solution
Arrhenius base	a substance whose water (aqueous) solution contains or yields/produces <b>OH<sup>-</sup> (HYDROXIDE IONS)</b> as the <b>ONLY NEGATIVE ION</b> when dissolved in water.
Bronsted-Lowry acid	Proton Donor
Bronsted-Lowry base	Proton Acceptor
Electrolyte	substances that conduct electric current when dissolved in water
hydronium ion	H <sub>3</sub> O <sup>+</sup> ION
hydroxide ion	OH <sup>-</sup> ION
indicator (acid/base)	substance that <b>CHANGES COLOR</b> as a result of a <b>pH CHANGE</b>
neutralization	when an Arrhenius acid and an Arrhenius base react to form <b>WATER</b> and a <b>SALT</b>
pH scale	designed to measure <b>HOW ACIDIC</b> or <b>HOW BASIC</b> an aqueous solution is
titration	used to <b>CALCULATE THE CONCENTRATION (MOLARITY) OF AN UNKNOWN SOLUTION</b>