Term	Definition
Anode	SPONTANEOUSLY LOSES ELECTRONS to cathode; the NEGATIVE electrode in a VOLTAIC CELL; electrode where OXIDATION occurs; The MORE ACTIVE of the 2 metals goes here
Cathode	SPONTANEOUSLY ATTRACTS ELECTRONS to it; the POSITIVE electrode in a VOLTAIC CELL; electrode where REDUCTION occurs; The LESS ACTIVE of the 2 metals goes here
Electrochemical Cell	Involve redox reacions and the flow of electrical energy
Electrode	(conductive surfaces where oxidation or reduction occurs
Electrolytic Cell	Cells that use ELECTRICAL ENERGY to force a NONSPONTANEOUS CHEMICAL REACTION to occur
Galvanic or Voltaic Cell	Cells that SPONTANEOUSLY convert CHEMICAL energy into ELECTRICAL energy or electric CURRENT; BATTERIES
Half reaction	allow us to show the EXCHANGE OF ELECTRONS in a redox rxn; 1 for oxidation and 1 for reduction
Oxidation Half Reaction	Loss OF ELECTRONS by an atom or ion
Oxidation Number	POSITIVE, NEGATIVE, OR NEUTRAL (ZERO) VALUES that can be assigned to atoms; used to identify how many electrons are being lost or gained by an atom/ion when they FORM BONDS
Oxidizing Agent	SPECIES that is REDUCED; species that DOES THE OXIDIZING
Redox	reactions that involve the TRANSFER OF ELECTRONS
Reducing Agent	SPECIES that is OXIDIZED; species that DOES THE REDUCING
Reduction Half Reaction	GAIN OF ELECTRONS by an atom or ion
Salt bridge	provides a path for the FLOW OF IONS between the half-cells